

IPC Cycle Sheet

March 30, 2020 thru April 3, 2020

Goals: TLW learn and demonstrate how to write formulas and name compounds.

Monday: Watch the Chemical Formulas Videos

Homework: Do worksheet
Chemical Formulas

Tuesday: Watch the Writing Ionic Formulas Video
Watch the Writing Covalent Formulas Video

Homework: Warm Up #66 and 67

Wednesday: Watch the Chemical Naming Videos

Homework: Do worksheet
Naming Compounds

Thursday: Class @9:00 AM

Homework: Check your work (score keys)
Warm Up #61 and 62

Friday: Watch Naming Ionic Compounds Video
Watch Naming Molecular Compounds Video

Homework: Warm Up #63 and 64

Vocabulary

anion	ionic bond	polyatomic ion
cation	covalent bond	chemical bond
salt	metallic bond	chemical symbol
molecule	polar/non polar	chemical formula

Know the following

chemical bonding	predicting bonds
oxidation numbers	determining charge
prefixes	polyatomic ions
element roots	

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Integrated Physics and Chemistry



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Chemical Symbol

A symbolic representation of the elements.

Examples: gold Au
silver Ag
hydrogen H

Binary Compound

A binary compound consists of two parts, usually a cation and an anion.

Keys to Chemical Formulas

The keys to writing and naming chemicals are learning to use the periodic table, knowing the polyatomic ions and knowing the prefixes.

Chemical Formula

A symbolic representation of a chemical substance.

Examples: NaOH is sodium hydroxide

H₂O is water

Polyatomic Ion

A covalently bonded group of atoms that have a positive or negative charge.

Chemical Reaction

It's the process in which substances undergo physical and chemical changes to produce new substances.

Reactants

Reactants are the starting substances in a chemical reaction.

Products

Products are the new substances formed in a chemical reaction.

Conservation of Mass

Mass can neither be created nor destroyed.

Reaction Types

synthesis – combination of two or more substances to create one more complex substance.

decomposition – the breaking down of a complex substance into two or more simple substances.

single replacement – uncombined element replaces an element that is part of a compound.

double displacement – different atoms in two different compounds replace each other

combustion – a substance reacts with oxygen to produce water, CO₂ and heat.

Mole

A mole is an Avogadro's number of anything.

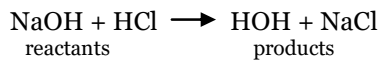
Avogadro's Number

Avogadro's number is the number of atoms in exactly 12.00 grams of carbon-12.

$$N = 6.022 \times 10^{23}$$

Chemical Equation

A chemical equation is a **symbolic description** of a chemical reaction.



It tells you 3 things:

1. The reactants
2. The products
3. The mole ratios (amounts)

Symbols

(g) = gas

(l) = liquid

(s) = solid

(aq) = aqueous

→ = yields

Δ = change

+ = combine

↑ = released